

Name: Key

Period: _____ Table: _____

/ 10

The Nuclide Symbol of an atom is given.

- Identify the Number of Protons, Neutrons and Electrons in the spaces provided
- Write the Electron Configuration
- Add the electrons to the Bohr Diagram

Nuclide
Symbol

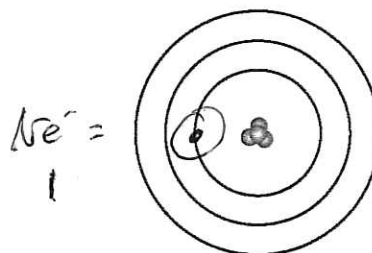
a) p^+ , n^0 & e^-

c) Bohr Diagram

1.

 ${}^3_1\text{H}$

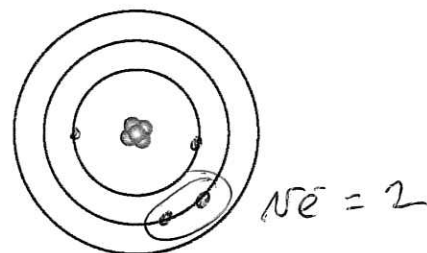
1 # of protons
2 # of neutrons
1 # of electrons

b) Electron Configuration: $1s^1$


2.

 ${}^{10}_4\text{Be}$

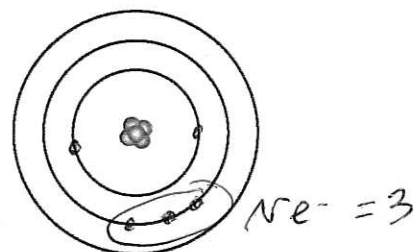
4 # of protons
6 # of neutrons
4 # of electrons

b) Electron Configuration: $1s^2 2s^2$


3.

 ${}^{11}_5\text{B}$

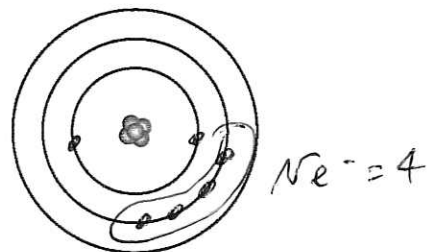
5 # of protons
6 # of neutrons
5 # of electrons

b) Electron Configuration: $1s^2 2s^2 2p^1$


4.

 ${}^{14}_6\text{C}$

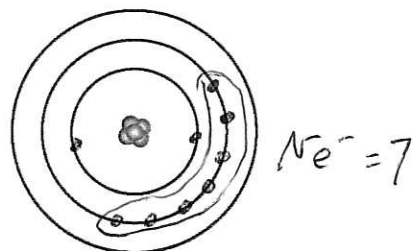
6 # of protons
8 # of neutrons
6 # of electrons

b) Electron Configuration: $1s^2 2s^2 2p^2$


5.

 ${}^{20}_9\text{F}$

9 # of protons
11 # of neutrons
9 # of electrons

b) Electron Configuration: $1s^2 2s^2 2p^5$


Nuclide
Symbol

a) p^+ , n^0 & e^-

c) Bohr Diagram

6.

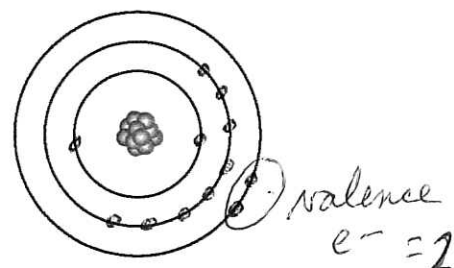
$^{24}_{12}\text{Mg}$

12 # of protons

12 # of neutrons

12 # of electrons

b) Electron Configuration: $1s^2 2s^2 2p^6 3s^2$



7.

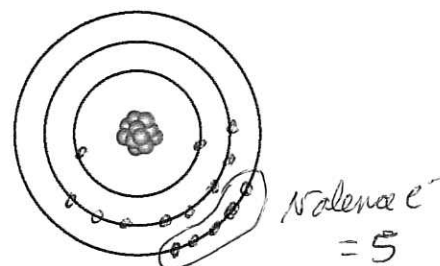
$^{32}_{15}\text{P}$

15 # of protons

17 # of neutrons

15 # of electrons

b) Electron Configuration: $1s^2 2s^2 2p^6 3s^2 3p^3$



8.

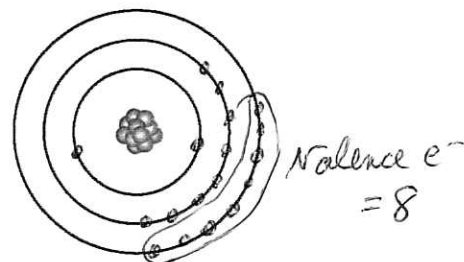
$^{40}_{18}\text{Ar}$

18 # of protons

22 # of neutrons

18 # of electrons

b) Electron Configuration: $1s^2 2s^2 2p^6 3s^2 3p^6$



9.

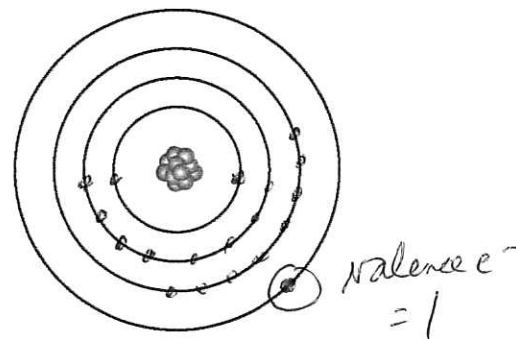
$^{39}_{19}\text{K}$

19 # of protons

20 # of neutrons

19 # of electrons

b) Electron Configuration: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$



10.

$^{40}_{20}\text{Ca}$

20 # of protons

20 # of neutrons

20 # of electrons

b) Electron Configuration: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$

